

Name _____

Date _____

Due Date _____

KEY

Mark _____ /34

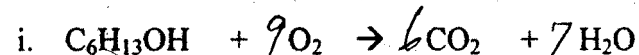
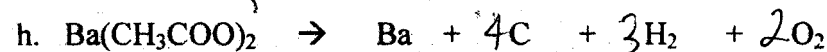
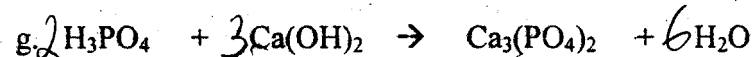
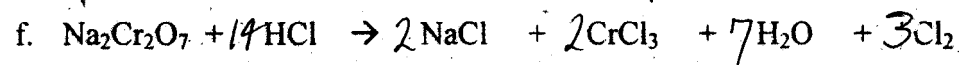
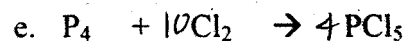
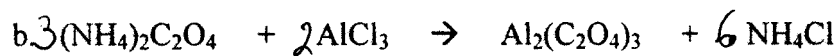
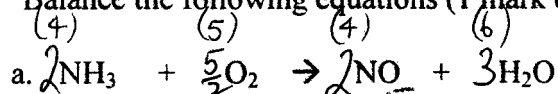
Correct and Hand in Again by _____

Chemistry 11

Hand In Assignment # 7—Chemical Equations

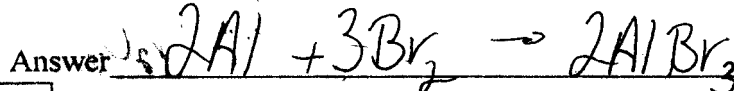
This Assignment will be marked and you are allowed to do one set of corrections.

1. Balance the following equations (1 mark each = 10 marks)



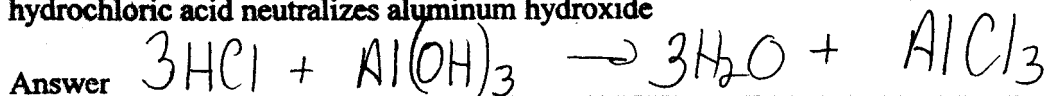
2. Write a balanced chemical equation for each of the following. Don't forget *diatomic* elements! (2 marks each = 24 marks)

- a. aluminum metal reacts with bromine to form aluminum bromide

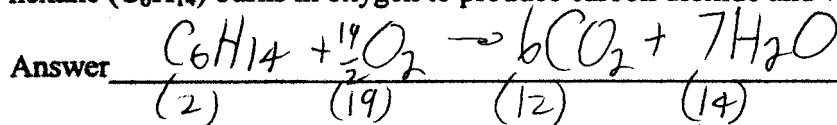


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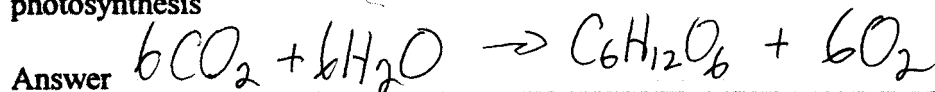
b. hydrochloric acid neutralizes aluminum hydroxide



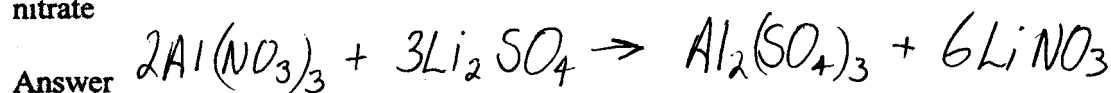
c. hexane (C_6H_{14}) burns in oxygen to produce carbon dioxide and water



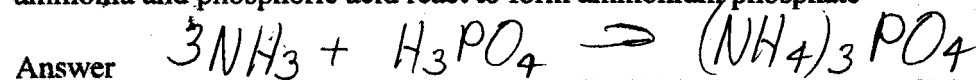
d. carbon dioxide and water are reacted to produce glucose ($\text{C}_6\text{H}_{12}\text{O}_6$) and oxygen in photosynthesis



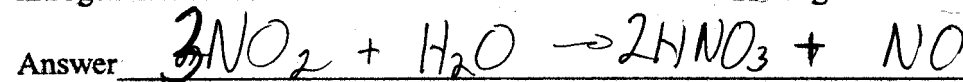
e. aluminum nitrate reacts with lithium sulphate to form aluminum sulphate and lithium nitrate



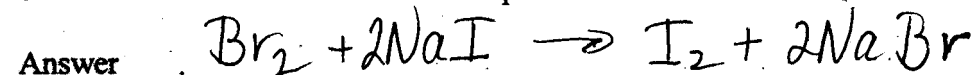
f. ammonia and phosphoric acid react to form ammonium phosphate



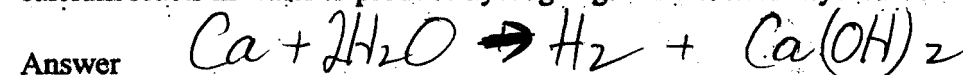
g. nitrogen dioxide reacts with water to form nitric acid and nitrogen monoxide



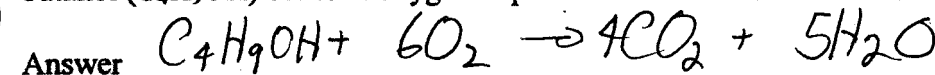
h. bromine reacts with sodium iodide to produce iodine and sodium bromide



i. calcium reacts in water to produce hydrogen gas and calcium hydroxide



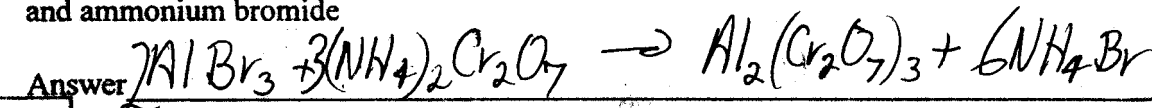
j. butanol ($\text{C}_4\text{H}_9\text{OH}$) burns in oxygen to produce carbon dioxide and water



k. hydrogen peroxide decomposes to form water and oxygen gas



l. aluminum bromide reacts with ammonium dichromate to produce aluminum dichromate and ammonium bromide



$\frac{22}{22}$